

Chemical Formulas Key Term Review

PART A Decide whether the underlined term in each sentence is used correctly. If so, write *true* in the space provided. If not, change the underlined term to a term in the box.

| | | |
|------------------|-------------------|----------------|
| binary compound | formula mass | subscript |
| chemical formula | oxidation number | valence |
| electrons | diatomic molecule | polyatomic ion |

- _____ 1. A subscript is the sum of the mass numbers of all the atoms in a molecule.
 _____ 2. A chemical formula is a way of writing the name of a compound using chemical symbols.
 _____ 3. A polyatomic ion is a molecule made up of two atoms of the same element.
 _____ 4. Formula mass is the number of electrons an atom gains, loses, or shares when it forms a chemical bond.
 _____ 5. Electrons in the outermost energy level of an atom are called valence electrons.

PART B Circle the term in each group that does not belong with the others. Then, in the space provided, explain how the terms in the remaining pair are related.

1. diatomic molecule, element, oxidation number _____

2. polyatomic ion, subscript, chemical symbols _____

3. binary compound, mass number, formula mass _____

4. outermost energy level, chemical formula, valence electrons _____

5. chemical bond, subscript, oxidation number _____
