

## Writing Chemical Formulas

### Enrichment Activity

**Skills:** *applying, interpreting*

Write a balanced equation for each of the chemical reactions described below. Remember that most gaseous elements form diatomic molecules. Use your text if you need help with formulas or symbols.

1. Carbon plus oxygen yields carbon dioxide. \_\_\_\_\_  
\_\_\_\_\_
2. When iron and sulfur are heated, they react to form a new compound called iron sulfide.  
\_\_\_\_\_
3. Hydrochloric acid is formed in a reaction between hydrogen gas and chlorine gas. \_\_\_\_\_  
\_\_\_\_\_
4. Lead and copper(II) nitrate react to produce copper and lead(II) nitrate. \_\_\_\_\_  
\_\_\_\_\_
5. When zinc reacts with hydrochloric acid, zinc chloride and hydrogen gas are produced. (Zinc has an oxidation number of 2+.) \_\_\_\_\_  
\_\_\_\_\_
6. When methane ( $\text{CH}_4$ ) is burned, the products are carbon dioxide and water. \_\_\_\_\_  
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